

NCERT SOLUTIONS

CLASS-XII CHEMISTRY

CHAPTER-15

POLYMERS

Q1: Define the following terms: monomer and polymer.

Ans:

Polymers can be defined as a macro-molecules having high molecular mass, which are made of repeating structural units derived from monomers. Polymers consists of a high molecular mass ($10^3 - 10^7 u$). In one polymer, numerous monomer units are combined together by strong covalent bonds. Polymers can be both synthetic and natural. Rubber, polythene and nylon 6, 6 are examples of polymers. Simple and reactive molecules called as monomers, fuse with each other in large groups through covalent bonds to give rise to polymers. **Eg: propene, ethene, vinyl chloride, styrene.**

Q2: Explain natural and synthetic polymers with the help of two examples of each type?

Ans:

Polymers that are found in nature are called Natural Polymer. They are formed from plants and animals. Various examples of natural polymer are protein, starch, cellulose, etc

Polymers made by human beings are known as Synthetic Polymer. **Various examples of synthetic polymer are synthetic rubbers (Buna-5), plastic (polythene), synthetic fibres (nylon 6, 6).**

Q3: Differentiate between homo-polymer and co-polymer by giving example of each.

Ans

HOMO-POLYMER	CO-POLYMER
Polymerization of a single monomer leads to formation of polymers which are known as homo-polymers. In other words, from one monomer, the repeating units of homo-polymers are formed. Example: homo-polymer of ethane is polythene.	Co-polymers are polymers whose repeating units are derived from monomers which are of two types. Examples: Co-polymer of 1, 3 - butadiene and styrene is Buna - S.

Q4: Briefly explain the functional of a monomer?

Ans:

The functional of a monomer may be defined as the total number of binding sites, which are present in that particular monomer.

For example, adipic acid and 1, 3-butadiene is two and that of propene and ethene is one

Q5: Explain the term polymerization.

Ans:

The process of forming high molecular mass ($10^3 - 10^7 u$) macro-molecules, consisting of repeated structural units formed or derived from monomers is known as **polymerization**. Various monomer units are joined by strong covalent bonds in a polymer.

Q6: Determine whether $(-NH-CHR-CO-)_n$, is a homo-polymer or a co-polymer?

Ans:

$(-NH-CHR-CO-)_n$ is a **homo-polymer**, the reason being that it is derived from a single monomer unit, $NH_2-CHR-COOH$.

Q7: Determine the classes in which the polymers are classified on the basis of molecular forces?

Ans:

On the basis of inter molecular forces magnitude present in polymers, polymers are classified in groups given below:

(i) Elastomers

(ii) Fibres

(iii) Thermosetting polymers

(iv) Thermoplastic polymers

Q8: Differentiate between condensation and addition polymerization?

Ans:

Condensation Polymerisation :The process in which the polymers are formed by the repeating condensation reactions in between the two different bifunctional or trifunctional monomers. In this process molecules such as hydrochloric acid or water is eliminated.

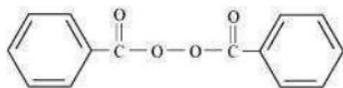
For example, nylon 6, 6 is the result of condensation polymerization of hexamethylenediamine and adipic acid.

- (1) Vinyl chloride ($CH_2 = CHCl$)
 (2) Tetrafluoroethylene ($CF_2 = CF_2$)
 (3) Formaldehyde (HCHO) and phenol (C_6H_5OH)

Q 13: Write the name and structure of one common initiators, which are used in free radical addition polymerisation.

Ans:

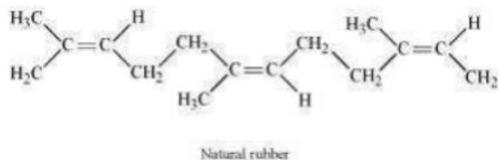
Benzoyl peroxide is one of the common initiator used in free radical addition polymerization.



Q 14: How do the presence of double bonds in rubber molecules influence their structure and reactivity?

Ans:

Natural rubber is a linear cis-polyisoprene in which the double bonds are present between C_2 and C_3 of the isoprene units.



Because of this cis-configuration, intermolecular interactions between the various strands of isoprene are quite weak. As a result, various strands in natural rubber are arranged randomly. Hence, it shows elasticity.

Q 15: Discuss the main purpose of vulcanization of rubber.

Ans:

Natural rubber is useful but has problems associated with its use. The disadvantages of natural rubber are as follows:

1. Natural rubber is sticky and soft at room temperature. At elevated temperatures i.e. greater than 335 K, it becomes even softer. At low temperatures i.e. less than 283K, it becomes brittle. Therefore natural rubber can be used only at temperature range of 283 K-335 K to maintain its elasticity.
2. It absorbs large amount of water
3. It has low resistance to abrasion and low tensile strength.
4. It is soluble in non-polar solvents.
5. Can be easily attacked by oxidizing agents.

Vulcanization is done mainly to improve the properties of natural rubber. In this process, a mixture of raw rubber with sulphur and appropriate additive is heated at a temperature range between 373 K and 415 K.

Q 16: Give the monomeric repeating units of Nylon-6 and Nylon-6, 6?

Ans:

The monomeric repeating unit of nylon 6 is $[NH - (CH_2)_5 - CO]$, which is derived from Caprolactam.

The monomeric repeating unit of nylon 6, 6 is $[NH - (CH_2)_6 - NH - CO - (CH_2)_4 - CO]$, which is derived from hexamethylene diamine and adipic acid.

Q 17: List out the names and its structures of the monomers for the following polymers:

- (i) Buna-S (ii) Buna-N
 (iii) Dacron (iv) Neoprene

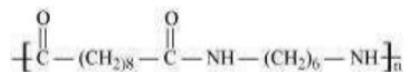
Answer

Polymer	Monomer	Structure of monomer
i Buna-S	1, 3-butadiene	$CH_2 = CH - CH = CH_2$
	Styrene	$C_6H_5CH = CH_2$
ii Buna-N	1, 3-butadiene	$CH_2 = CH - CH = CH_2$
	Acrylonitrile	$CH_2 = CH - CN$
iii Neoprene	Chloroprene	$CH_2 = \overset{Cl}{\underset{ }{C}} - CH = CH_2$
iv Dacron	Ethylene glycol	$HOH_2C - CH_2OH$

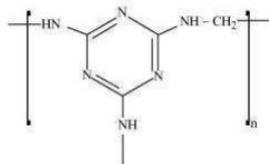


Q 18: Identify the monomer in the following polymeric structures.

(i)



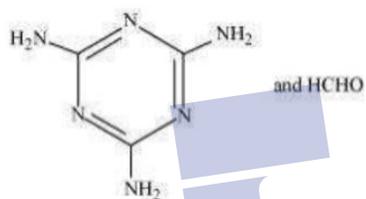
(ii)



Ans:

(i) The monomers of the given polymeric structure are hexamethylene diamine $[H_2N(CH_2)_6NH_2]$ and decanoic acid $[HOOC - (CH_2)_8 - COOH]$.

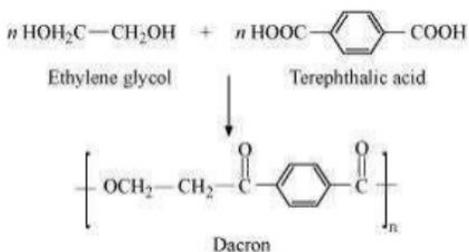
(ii) The monomers of the given polymeric structure are:



Q 19: How can dacron be obtained from terephthalic acid and ethylene glycol?

Ans:

Dacron is formed by the condensation polymerisation of terephthalic acid and ethylene glycol.



Q 20: What is a biodegradable polymer? Give an example of a biodegradable aliphatic polyester.

Ans:

A polymer, which can be decomposed by bacteria are called as Biodegradable polymer. Poly - β - hydroxybutyrate - CO - β - hydroxyvalerate (PHBV) an example of a biodegradable aliphatic polyester.

